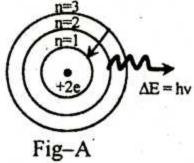
Model Question of HSC Examination 2017 (All Board)

Sub: Chemistry 1st paper (Creative)

Time: 2 Hrs 10 min [Answer any four questions]	Sub Code : 1 7 6 Full marks: 4
 A chemist wnats to produce ammonia Nitrogen & hydrogen gas in a closed vasse substance B which does no take part in chemica a. What is formula of olium b. What is endothermic reaction? Give an exact. c. Describle the effect of temperature production of mentioned gas. 	al and mixed a al reaction. 1 mple. 2
d. Write down the classification of substance Air Air Perforate vineage to bottom	_
 a. What is quagulation? b. Why caustic soda is not used as glass cleaned. c. Write down the importance of mentioned preserving foods d. How you get this product from Sugar-Explains. 3. 	d chemicals in



a. What is common ion effect?

b.	Who mominated electron? Find out the charge of it 2				
C	Write down the postulates of Fig A				
d.	Explain hydorgen spectrum from the stem. 4				
4.					
E	lement→ Na Mg Al Si P S Cl Ar				
a.	What is the bond angle of ∠HOH in water molecule? 1				
b.	What is d block element? Write down the electronic				
	configuration of ther outermost shell				
c.	Electro negativity is a periodic property-Explain according				
	to stem.				
d.	Discouss Acid base properties of oxides of memtioned				
	element 4				
5.	► PCl ₅ ↔ PCl ₃ + Cl ₂				
a.	What is chemical equilibrium?				
b.	Explain the dynamism of chemical equilibrium 2				
c.					
	explain it in the light of stem.				
d.	Reactant is decomposed 25% at 70°C. Total pressure				
	1.75atm. Find out Kp & Kc. 4				
6.	▶ In chemistry lab a student wants some green FeSO ₄ but				
	en she found she noliced that its color is not green. She tried				
	know the reaction.				
a.	What is electro negagivity of hydrogen.				
	Why all d-block elemetrs are not transition elements? 2				
	"Metal of the compound form complex compound"-				
po.	Explain with your logic.				
d,	Show another color compound & give the reason of its				
4	color. 4				

Model Question of HSC Examination 2017 (All Board)

Sub - Chemistry (MCQ)

Sub Code : 1 7 6 Full Marks: 35

1	N.B. Fill the circle of the correct answer with a	black ball	point	pen. Eacl	question	bears i	mark.
	****	a	Carl	hovilio oc	id group		7/12/3

- Which compound may cause explosion on mixing with air?
 - (a) Ozonide
- (b) Methanol
- © Nitric acid
- Copper sulphate
- 2. A quanta is

Time: 35 Minutes

- Matter can absorb a certain amount of
- (b) Matter can release large amount of
- © Matter emits energy discontinuously
- @ Matter does not absorb or emit energy.
- Bohr's atomic model is related to
 - (a) Wave of electron (b) Quanta mechanics
 - © H-Spectrum
- d mechanics
- Principle quantum number determines-
 - The shape of the orbital
 - (b) the number of possible orientation
 - © The size of the orbital
 - d the spin direction
- 5. When principle quantum number n = 4, then orbital number and electron number
 - (a) 9.18
- **ⓑ** 10, 20
- © 15,30
- @ 16,32
- 6. Entry of electrons in energy level is related to
 - Hund's rule
 - (b) Aufbau rule
 - © Pauli's Exclusion rule
 - (d) wave function
- A-Sub group of a element is determined by
 - (a) Electron enters f-orbital
 - (b) Electron enters into d-orbital
 - © Eiement has no d-orbital
 - Element has partial fulfill d and f orbital
- 8. What is the wavelength of green color?
 - (a) 424-491nm
- (b) 575-585 nm
- © 400-424 nm
- @ 491-575 nm
- 9. Wave number 2750-2700 cm⁻¹ which indicates?
 - Keto group
- Aldehyde group
- Alcohol group

- Carboxilic acid group
- 10. Solubility product is related to
 - Concentration of ions
 - (b) partial pressure of ions
 - © solvent types @ aqueous solvent
- 11. Which is the correct order of orbital energy?
 - 3s < 3d < 4p</p>
- 3d<3s<4p
 </p>
- @ 4p<3s<3d
- @ 3s<4p<3d
- 12. Sc and Zn are
 - a d-block and transition elements
 - (b) transition elements
 - © Inner transition elements
 - d-bolck elements
- 13. Melting point and boiling point increases when
 - a Nuclear charge decreases
 - Nuclear charge is absent
 - © Nuclear charge increases
 - Muclear charge is present
- 14. Polarization is related to
 - (a) Covalent character
 - (b) Ionic character
 - © Coordinate covalent character
 - d non-polar molecule
- 15. How many bonds are present in tetramin cupric sulphate?
 - (a) 2
- (b) 3
- (c) 4
- (d) 5.
- 16. Which one is correct?
 - a MgCl₂ shows more covalent character than AlCla
 - AlCl₃ shows more covalent character than MgCl₂
 - © AClCl₃ has high melting point MgCl₂
 - (d) AlCl₂ show more ionic character.
- 17. Bond becomes strong when
 - a electron occupy the volume of outside of two atoms.
 - b electron occupy the place between two nucleus of atoms
 - © partial overlap of atomic orbital
 - (d) concentration of electron is low.

 ⑤ head to head overlap of orbital ⑥ donating of electrons ⑤ partial overlap of atomic orbital ⑥ partial overlap of hybrid orbital. 19. What will be the hybridization of centre catom of (Fe(CN)₃)²? ⑥ sp³d	18. Oxygen molecule is formed by	© diastase @ maltase
 ⊕ donating of electrons ⊕ partial overlap of atomic orbital ⊕ partial overlap of hybrid orbital. 19. What will be the hybridization of centre atom of (Fe(CN)_a)²? ⊕ sp³d² ⊕ HNO₁ is less stronger acid that HNO₂ ⊕ HNO₁ is strue? ⊕ HNO₁ is strue? ⊕ Both are equal stronger. 28. At 30° and 1.5 atm. Pressure 15.6%PCls dissociates, Claculate the value of Kp at this temperature. ⊕ 3.74 × 10² atm ⊕ 3.0 × 10³ atm ≥ 0.0 × 10³ atm ≥ 2.0 × 10³ atm ≥ 2.0 × 10³ atm ⊕ 2.0 × 10³ atm ⊕ 2.0 × 10³ atm ⊕ Sp³d² ⊕ Production is less 30. The dissociation constant of 1 molar solution of HCN is 4 × 10⁻ activation energy and rate constant with temperature ⊕ activations energy and rate constant with temperature ⊕ rate constant with temperature ⊕ rate of reaction ⊕ Rate is decreased ⊕ production is less 30. The dissociation constant of 1 molar solution of HCN is 4 × 10⁻ activation energy is produced from ⊕		26. Hess's law deals with
© partial overlap of atomic orbital ⊕ partial overlap of hybrid orbital. 19. What will be the hybridization of centre atom of (Fe(CN) ₆] ²⁻ ? ② sp³d ⊕ sp³d² ② sp³d⁴ ⊕ sp³d² ② sp³d² ② sp³d⁴ ⊕ sp³d² ② sp³d² ② sp³d² ⊕ sp³d² ② sp³d² ③ sp³d² ② sp³d² ② sp³d² ③ sp³d² ② sp³d		Changes in heats of reaction
(a) partial overlap of hybrid orbital. 19. What will be the hybridization of centre atom of (Fe(CN _b) ¹ -? (a) sp ³ d (b) sp ³ d ² (b) sp ³ d (c) sp ³ d ² (c) sp ³ d (d) sp ³ d ² (e) sp ³ d (d) sp ³ d ² (d) sp ³ d (d) s		Profession of the State of the Control of the State of th
19. What will be the hybridization of centre atom of (Fe(CN) _a) ² ? ② sp³d ③ sp²d² ② sp³d ⑤ sp³d² 20. Why bond angle of water is 104° instead of 109° ③ Lone pair-lone pair attraction force is less than long pair-bond pair ⑤ Bond pair-bond pair attraction force is high than lone pair-bond pair ⑥ None of them ② Lone pair-lone pair attraction force is high than lone pair-bond pair ⑥ None of them ② Lone pair-lone pair attraction force is high than lone pair-bond pair ⑥ None of them ② Lone pair-lone pair attraction force is high than lone pair-bond pair ⑥ None of them ② Lotahedral ⑥ Trigonal pyramidal ⑥ Bipyramidal ⑥ Octahedral ⑥ Square palanar ② Arrhenius equation relates ⑥ activation energy with temperature ⑥ rate constant with temperature ⑥ activations energy and rate constant with temperature ② Arrhenius parameter with temperature ③ Arrhenius parameter with temperature ② Collision between two molecules with less energy ② Collision between two molecules with less energy ② Collision between two melecules with less energy ② Collision with definite direction and sufficient energy ② Collision with definite direction and sufficient energy ③ Collision with definite direction and sufficient energy ⑤ Collision with definite direction and sufficient energy ⑥ Collision with defini		© equilibrium constant
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Sp'd*		
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© Octahedral ② Square palanar 22. Arrhenius equation relates ③ activation energy with temperature ⑤ rate constant with temperature ⑥ activations energy and rate constant with temperature ② activations energy and rate constant of 1 molar solution of HCN is 4 × 10 ⁻¹⁰ Calculate the percentate of molecules dissociated. ② 2.10 × 10 ⁻² % ⑤ 2.0 × 10 ⁻³ % ③ 2.10 × 10 ⁻³ % ④ 3.0 × 10 ⁻³ % ③ Endothermic ⑥ Exothermic ⑥ Reversible ② Irriversible ③ Alcohol ⑥ Acetic acid ② propanoic acid ④ Sodium benzoate ③ Glucose ⑥ Amino acid ② Aldehyde ④ Ketone ③ Glucose ⑥ Amino acid ② Phosphorus. 34. Which is additive? ③ Sulphur acid ⑥ Sulphur ② Suphonic acid ④ Phosphorus. 35. Which substance is used glass cleaner? ④ Dye ⑥ Sulphur ② Carboxilic acid ④ Phosphoric acid		(1) 1 (1)
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22. Arrhenius equation relates a activation energy with temperature b rate constant with temperature c activations energy and rate constant with temperature d Arrhenius parameter with temperature 23. When the surface area increases reaction rate a not changes b decreases c increases d changes 24. Reaction happens when Collision between two molecules with less energy b Collision between two melecules with activation energy Collision with definite direction and sufficient energy d come in contact with each other. 25. Starch is converted to maltose in action of Urease b Zymase 30. The dissociation constant of 1 molar solution of HCN is 4 × 10 ⁻¹⁰ Calculate the percentate of molecules dissociated. a 2.10 × 10 ⁻²⁰ b 2.0 × 10 ⁻²⁰ c 2.0 × 10 ⁻³ w d 3.0 × 10 ⁻³		
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